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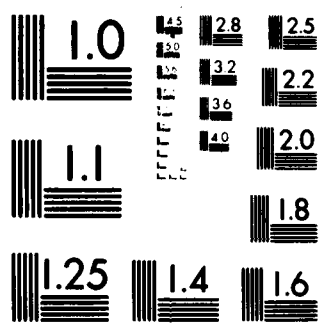
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SIDE GROUP CONSTRUCTION IN HIGH POLYMERIC PHOSPHAZENES VIA LITH--ETC(U)  
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SIDE GROUP CONSTRUCTION IN HIGH POLYMERIC  
PHOSPHAZENES VIA LITHIOPHENOXY INTERMEDIATES

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Side Group Construction in High Polymeric Phosphazenes  
via Lithiophenoxy Intermediates

H. R. Allcock,\* T. J. Fuller, and T. L. Evans

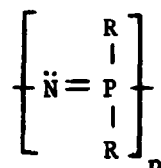
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**ABSTRACT** The new lithiophenoxyphosphazene high polymers,  $[\text{NP}(\text{OC}_6\text{H}_4\text{Li-p})_x(\text{OC}_6\text{H}_4\text{Br})_y]_n$  and  $[\text{NP}(\text{OC}_6\text{H}_4\text{Li-p})_x(\text{OPh})_y]_n$  have been prepared by a metal-halogen exchange reaction between *n*-butyllithium and  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_2]_n$  or  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_x(\text{OPh})_y]_n$  at  $-40^\circ\text{C}$  to  $-60^\circ\text{C}$ . The lithio-derivatives were used as reactive intermediates for the synthesis of polymers of formula:  $[\text{NP}(\text{OC}_6\text{H}_4\text{X-p})_x(\text{OC}_6\text{H}_4\text{Br-p})_y]_n$ , where  $\text{X} = -\text{PPh}_2$ ,  $-\text{PPh}_2\text{C}_4\text{H}_9^+\text{Br}^-$ ; or  $[\text{NP}(\text{OC}_6\text{H}_4\text{X-p})_x(\text{OPh})_y]_n$ , where  $\text{X} = -\text{PPh}_2$ ,  $-\text{PPh}_2\text{C}_4\text{H}_9^+\text{Br}^-$ ,  $-\text{SnPh}_3$ ,  $-\text{AuPPh}_3$ , or  $-\text{COOH}$ . Evidence was obtained that the replacement of bromine by lithium is rapid up to the point at which one half of the phenoxy groups bear a lithium substituent. Beyond that stage lithiation is extremely slow. The replacement of lithium by the electrophiles,  $\text{ClPPh}_2$ ,  $\text{ClSnPh}_3$ ,  $\text{ClAuPPh}_3$ , or  $\text{CO}_2$  was rapid and complete. Quaternization of pendent  $-\text{PPh}_2$  groups by *n*-butyl bromide was also a facile reaction. The differences between these high polymeric reactions and those of analogous small molecule cyclic phosphazene model systems are discussed. Comparisons are also made with the lithiation reactions of poly(*p*-bromostyrene).

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The synthesis of macromolecules that can function as carrier molecules for catalysts or chemotherapeutic agents or which can selectively bind ions in aqueous media is a subject of growing importance. So also is the preparation of polymers that may serve as templates and supports for the construction of one-dimensional, electrically conducting arrays of metal atoms. We describe here an initial approach to possible solutions of these problems. This approach makes use of the polyphosphazene chain (I) as a carrier species.



I

The synthesis of a wide range of poly(organophosphazenes) by the interaction of the high polymeric reactive intermediates,  $(\text{NPCl}_2)_n$  (II) or  $(\text{NPF}_2)_n$ , with nucleophiles has been discussed in earlier papers.<sup>1-4</sup> In this paper we describe a new method for the attachment of organic or inorganic residues to polyphosphazene molecules. This synthetic pathway involves a two-step process with the initial attachment of an organic "spacer" side group to the inorganic skeleton, followed by subsequent construction reactions carried out on the organic unit. This route was developed for the following two reasons. First, it was necessary to exclude from consideration any synthetic pathway that involved the reaction of a difunctional reagent with  $(\text{NPCl}_2)_n$  or  $(\text{NPF}_2)_n$  because of the inevitable crosslinking that would ensue. The route developed in this work avoids this complication. Second, advantages can be foreseen

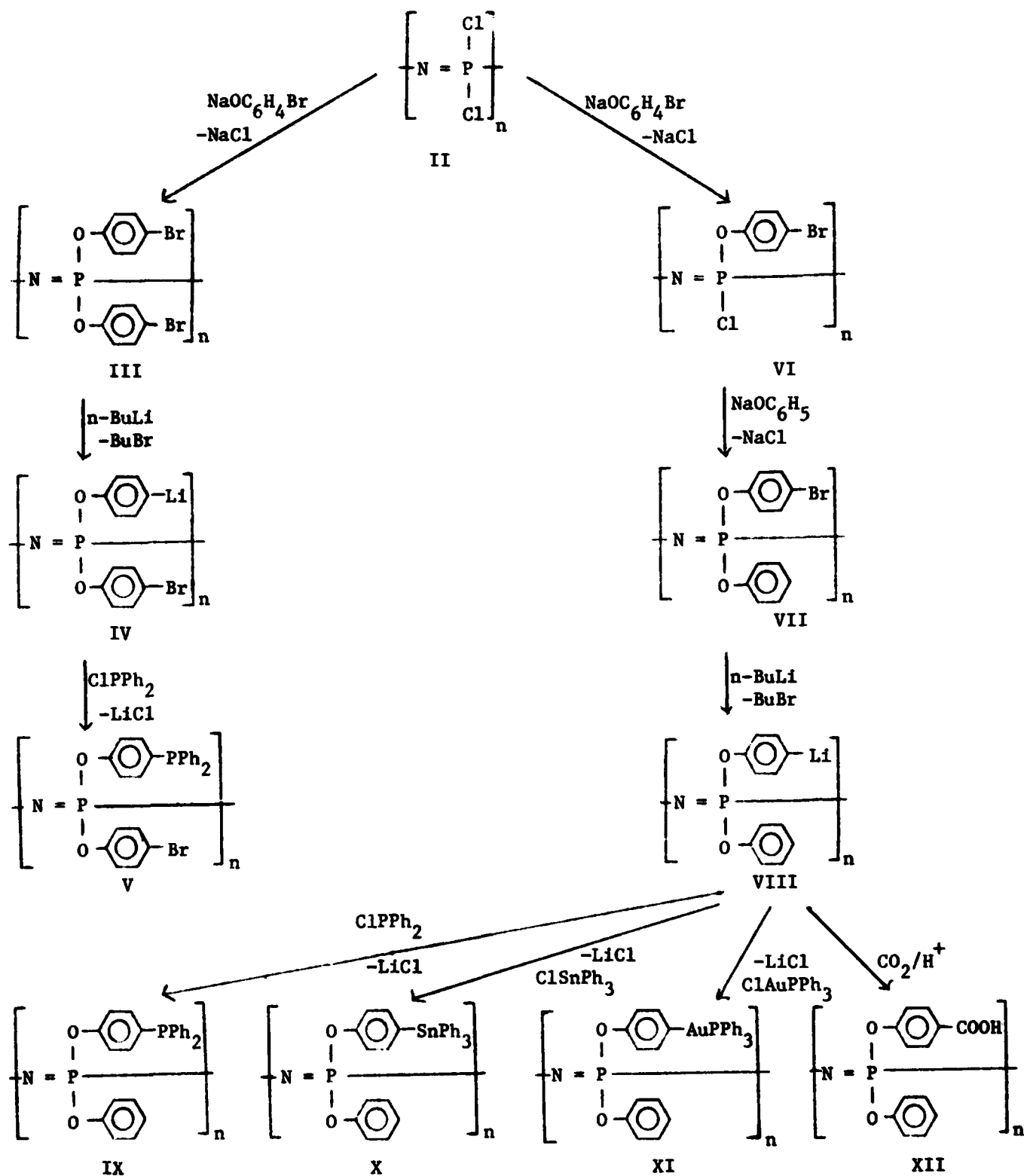
for the attachment of phosphines, alcohol functions, or carboxylic acid groups to the outer fringes of the side group structure. Such reactive functionalities could be employed to bind transition metals or could serve as sites for further chemical modification without interference from the polymer skeleton or from the side group spacer units.

The reaction sequence employed in this work is shown in Scheme 1. Exploratory model reactions to test the feasibility of this approach have been carried out with the cyclic phosphazene,  $(\text{NPCl}_2)_3$ .<sup>5</sup> In this paper we discuss the reactions shown in Scheme 1, together with the unexpected anomalies not encountered with the small molecule model system. The use of polymer III and its related derivative VII as substrates for lithiation reactions can be compared with the analogous reactions of poly(*p*-bromostyrene).

### Results and Discussion

General Reaction Sequence. The reaction sequences shown in Scheme 1 involve three hitherto unexplored steps: (1) the preparation of high molecular weight polyphosphazenes with the formulae  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_2]_n$  (III) or  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_x(\text{OC}_6\text{H}_5)_y]_n$  (VII); (2) replacement of the bromine atoms in III or VII by lithium with the use of a metal-halogen exchange reaction with *n*-butyllithium; (3) the reaction of lithiophenoxyphosphazenes of type IV or VIII with electrophiles such as diphenylchlorophosphine, triphenyltin chloride triphenylgold(I)chloride, or carbon dioxide to yield substituted derivatives such as V and IX-XII. These three aspects are discussed in turn.

Synthesis, Structure, and Properties of *p* Bromophenoxy-Substituted Polyphosphazenes. Polymers of type III were prepared by the interaction of poly(dichlorophosphazene) (II) with sodium *p*-bromophenoxide. Mixed substituent





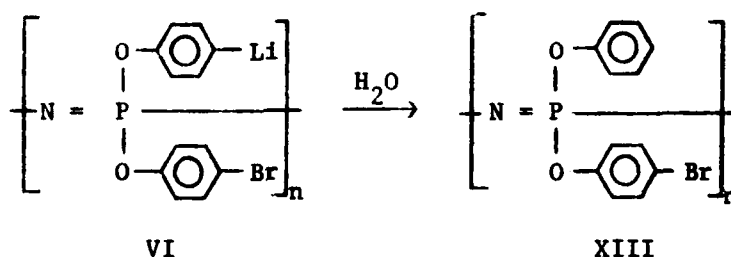
species, containing both *p*-bromophenoxy and phenoxy groups, were synthesized by the initial interaction of II with a stoichiometric deficiency of sodium *p*-bromophenoxide, followed by treatment with an excess of sodium phenoxide. Four different polymers corresponding to species of type VII were prepared. In these, the ratio of *p*-bromophenoxy to phenoxy groups was 0.1:1.9, 0.4:1.6, 0.7:1.3 and 1.12:0.88.

Evidence exists that the interaction of aryloxy ion with II results in a non-geminal substitution pathway.<sup>6</sup> Hence, the idealized non-geminal structure depicted in VII probably has some validity. The molecular weights and T<sub>g</sub> values for these polymers are listed in Table I. Comparisons of these polymers with samples of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_n$ , prepared under identical reaction conditions, showed that the *p*-bromophenoxy substituent group had little effect on the average molecular weight or the glass transition temperature. This is perhaps not surprising because the *p*-bromo functionality can exert very little additional steric hindrance either during substitution or during chain torsional motion. However, the solubilities of the polymers were changed by the presence of the bromine atoms. When 0%, 5%, 20%, or 35% of the side groups were *p*-bromophenoxy groups, the polymers were soluble in tetrahydrofuran (THF) (warming of the THF was often necessary to accomplish solubilization of the polymers), but polymers with 56% or 100% *p*-bromophenoxy groups were insoluble in THF unless they had been dissolved previously in boiling dioxane. It is conceivable that the *p*-bromophenoxy groups can generate a high degree of microcrystallinity which influences the solubility behavior.

The Metal-Halogen Reactions. The metal-halogen exchange reactions<sup>7</sup> were examined for the interaction at -40°C to -60°C in tetrahydrofuran of *n*-butyllithium with species of type VII or III that contained 5%, 20%, 35%, 56%

or 100% of the side groups as *p*-bromophenoxy units. The following features were found.

First, the possibility of side reactions during the lithiation process was monitored by treatment of the lithiophosphazene with water to yield products such as XIII. The composition of the products was then established by elemental microanalysis,  $^{31}\text{P}$  NMR spectroscopy, and by gel permeation chromatography (GPC) molecular weight determinations. The use of these techniques showed that neither chain scission reactions nor displacement of aryloxy side groups occurred to a detectable degree at  $-40^\circ\text{C}$  to  $-60^\circ\text{C}$ .



Second, by the use of elemental microanalysis, it was found that, although the complete lithiation reaction of the cyclic trimer,  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-}p)_2]_3$ , is rapid at  $-40^\circ\text{C}$  to  $-60^\circ\text{C}$ , the reaction of the polymeric analogue (III) is slower and proceeds in two stages. For example, the metal-halogen exchange process was rapid and nearly complete within 0.5 h for polymers of type VII in which 5%, 20%, or 35% of the side groups were bromophenoxy units. However, only partial lithiation took place in the same time period when the polymers contained 56% or 100% bromophenoxy groups. In fact, with these latter two polymers the metal-halogen exchange was rapid up to the point at which 50% of the total number of side groups bore a lithium atom. Beyond that point, the lithiation reaction was markedly retarded. As shown in Figure 1, which illustrates the effect of increasing reaction time on the lithiation of III

(followed by treatment with water to yield XIII and subsequent elemental microanalyses), lithiation was rapid in the first 0.5 h, during which time 51% of the bromine atoms had been replaced. Thereafter, the rate of lithiation was so slow that only 18% of the remaining bromine atoms were replaced in the next 5.5 h. The same behavior was shown by a polymer which initially contained 56% *p*-bromophenoxy groups. It appears likely that the presence of one lithiophenoxy unit inhibits the introduction of a second lithium atom at an aryloxy unit geminal to the first, perhaps by charge repulsion or possibly by aggregation of the aryllithium units. Evidence in favor of this interpretation was provided by the following experiment. Polymer III was allowed to undergo a prolonged reaction with *n*-butyllithium until 70% of the *p*-bromophenoxy groups had been converted to *p*-lithiophenoxy groups. Reaction of this compound with water yielded a product similar to XIII. Subsequent lithiation of this product was rapid and virtually complete. Hence, removal of the bound lithium component apparently removed the inhibiting factor.

A similar deactivation effect occurred during the metal-halogen exchange reaction between *n*-butyllithium and poly(*p*-bromostyrene). At -50°C in THF solvent, only 84% of the bromine atoms were replaced by lithium during 6 h.<sup>8</sup> Presumably the effect is less marked than with III because only one aryl group is attached to each skeletal linkage atom.

Crosslinking at 25°C. The reactions discussed in this paper require that the lithiation of III or VII and the subsequent reactions of IV or VIII must take place at low temperatures (-40°C to -60°C). At higher temperatures, crosslinking reactions take place. For example, if polymers of type VII (35% *p*-bromophenoxy-, 65% phenoxy-groups) were treated with *n*-butyllithium at -40°C and -60°C and the reaction mixture was treated with water at the same temperature,

uncrosslinked species of type XIII were formed. However, if the reaction mixture was allowed to warm to 25°C before treatment with water, the product was crosslinked and insoluble. It was shown that this crosslinking reaction was associated with the presence of aryl-lithium bonds, because treatment of the polymer,  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_n$ , with *n*-butyllithium at -50°C, followed by warming to 25°C and addition of water, did not yield a crosslinked polymer.

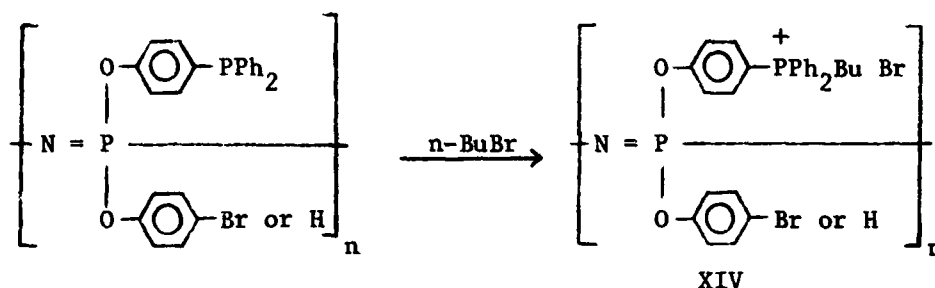
Lithiophenoxyphosphazene polymers could crosslink by a number of different mechanisms. First, intermolecular aryl coupling reactions could take place between *p*-lithiophenoxy and *p*-bromophenoxy substituent groups to form *p*-biphenyl-enedioxy crosslinks. Second, benzyne formation from *p*-bromophenoxy groups is possible, followed by coupling to *p*-lithiophenoxy residues. Third, a lithio-aryl unit could attack at phosphorus with displacement of a *p*-bromophenoxy residue and formation of a  $\text{P}-\text{O}-\text{C}_6\text{H}_4-\text{P}$  linkage. We favor this third mechanism. It has been reported previously<sup>9</sup> that the phenoxide groups can be displaced from the cyclic trimer,  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$ , by organolithium reagents at normal temperatures. Moreover, it was shown here that the treatment of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$  with phenyllithium yielded a mixture of products which contained the compound  $[\text{N}_3\text{P}_3(\text{C}_6\text{H}_5)(\text{OC}_6\text{H}_5)_5]$  (see Experimental section).

#### Interaction of Lithiophenoxypolyphosphazenes with Diphenylchlorophosphine.

Polymers of types VII or III with 5%, 20%, 35%, 56%, or 100% of the side groups as *p*-bromophenoxy units were subjected to lithium-halogen exchange reactions as described in the preceding sections. Treatment of the lithiophenoxy derivatives at -40°C to -60°C with diphenylchlorophosphine resulted in a high yield attachment of diphenylphosphorus residues to the phenoxy units to give products such as V or IX. The reaction conditions are summarized in Table II.

The coupling reactions with diphenylchlorophosphine appeared to be as efficient as those with water. The triarylphosphine unit is a valuable ligand for coordination to transition metal catalyst systems,<sup>10-12</sup> and this aspect will be discussed in a later paper.

The synthesis of V and IX was complicated by the concurrent formation of low percentages of butylphosphonium salts (XIV).



The degree of phosphonium salt formation in the conventional synthesis is listed in Table III. Presumably the quaternization reaction is a consequence of the liberation of *n*-butyl bromide during the metal-halogen process. The phosphonium salts were detected by three techniques. <sup>1</sup>H NMR spectroscopy was used to demonstrate the presence of butyl groups in the purified polymer. <sup>31</sup>P NMR spectra of a polymer of type IX, in which 35% of the aryloxy groups bore phosphorus-containing pendent groups, possessed a resonance at +24 ppm which was compatible with the presence of a phosphonium salt unit. This corresponded to approximately 3.1% of all the phosphine groups present (the backbone <sup>31</sup>P resonance was at -20 ppm and the phosphine resonance was at -8 ppm).<sup>13</sup> Moreover, partial proton decoupling experiments revealed that the resonance at +24 ppm was coupled to both aryl and alkyl protons. Finally, it was shown that the addition of an excess of butyl bromide to the polymer brought about a complete disappearance of the -8 ppm resonance and a concurrent growth in the size of the

peak at +24 ppm. Polymeric butylphosphonium salts are possible precursors for polymer-supported reactions<sup>14</sup> or for use as phase transfer catalysts.<sup>15</sup>

It is not obvious why the analogous cyclic model compound,  $[\text{NP}(\text{OC}_6\text{H}_4\text{P}(\text{OC}_6\text{H}_5)_2)_2]_3$ , was less prone to phosphonium salt formation during its synthesis. However, conversion of 50% of the phosphine groups to phosphonium units could be effected in the trimeric system by treatment with an excess of butyl bromide during 48 h at 80°C.

Interaction of Lithiophenoxyphosphazenes with Triphenyltinchloride, Triphenylphosphinegold(I)chloride or Carbon Dioxide. As shown in Scheme 1, lithio derivatives of type VIII were used for the attachment of triphenyltin, triphenylphosphinegold(I), or carboxylic acid functions to the aryloxyphosphazene structure, with the formation of X, XI, or XII. Specifically, polymers of formula  $[\text{NP}(\text{OC}_6\text{H}_4\text{SnPh}_3)_2(\text{OC}_6\text{H}_5)_{1.3}]_n$ ,  $[\text{NP}(\text{OC}_6\text{H}_4\text{AuPPh}_3)_2(\text{OC}_6\text{H}_5)_{1.48}]_n$  and  $[\text{NP}(\text{OC}_6\text{H}_5\text{COOH})_2(\text{OC}_6\text{H}_5)_{1.3}]_n$  were isolated in high yields. The details of the synthetic procedures are summarized in the Experimental section. The derivative,  $[\text{NP}(\text{OC}_6\text{H}_4\text{COOH})_2(\text{OC}_6\text{H}_4\text{Br})_{0.58}]_n$ , was also prepared in high yields from the interactions of IV with carbon dioxide.

#### Experimental Section

Equipment. <sup>31</sup>P NMR spectra were obtained in the Fourier transform mode at 40.4 MHz with the use of a JEOL PS-100 FT spectrometer and were processed with the use of a Nicolet 1080 computer. <sup>1</sup>H NMR spectra were obtained with the same spectrometer operated at 100 MHz. Infrared spectra were recorded with the use of a Perkin-Elmer model PE 580 high resolution infrared spectrophotometer. Approximate polymer molecular weights were determined with the

use of a Waters Associates AIC/GPC 501 instrument fitted with a 120 cm x 1 cm  $10^6$  styragel column.<sup>16</sup> Glass transition temperatures ( $T_g$ ) were measured with the use of a Chemical Instruments Corp. torsional braid analyzer kindly provided by N.A.S.A., Langley Field.

Materials. All experimental manipulations were performed under an atmosphere of dry nitrogen (Matheson). Tetrahydrofuran (THF) (Fisher) and dioxane (Fisher) were freshly distilled under nitrogen from sodium benzophenone ketyl. Poly(dichlorophosphazene),  $(\text{NPCl}_2)_n$ , was prepared from the melt-polymerization of  $(\text{NPCl}_2)_3$  (Ethyl Corporation) by methods described previously.<sup>17</sup> *n*-Butyllithium was used as received (Foote Mineral - 1.6 M solution in hexane). Phenol, *p*-bromophenol (Aldrich), diphenylchlorophosphine (Orgmet.), triphenyltinchloride (Strem), triphenylphosphinegold(I)chloride (Strem), and sodium hydride (Alfa) were used as received.

Elemental Microanalyses. The accuracy of elemental microanalyses, as determined by Galbraith Laboratories, was found to be dependent on the element that was analyzed and the composition of the polymeric compounds. Elemental analyses for bromine proved to be the most reliable indicator of polymer compositions. Furthermore, the polymers that possessed exclusively *p*-bromophenoxy substituent groups and those that contained only phenoxy and *p*-bromophenoxy substituent groups yielded the most accurate microanalytical data. The compositions of polymers that possessed phosphine residues could be determined by  $^{31}\text{P}$  NMR peak area integrations of the phosphazene and phosphine phosphorus resonances. The ratios of phosphazene:phosphine, as determined by  $^{31}\text{P}$  NMR methods, agreed to within 3% of the value as determined by elemental microanalyses for bromine.

General Procedure for the Synthesis of Polymers with the Formulae

$(\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_2)_n$ , III, and  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_x(\text{OC}_6\text{H}_5)_y]_n$ , VII. The specific reaction conditions used for these syntheses are listed in Table IV. The general experimental procedure was as follows. Sodium *p*-bromophenoxide was prepared by the addition of a solution of *p*-bromophenol in dioxane (100 mL of dioxane was usually employed) to a stirred suspension of a molar excess of sodium hydride and dioxane (100 mL). After 4 h, the reaction mixture was heated to reflux and was filtered. The filtrate was added to a stirred solution of poly(dichlorophosphazene) in dioxane (150 mL). The reaction mixture was stirred for 4 to 8 h at 25°C and was then added to a solution of sodium phenoxide, prepared in a similar manner to the preparation of sodium *p*-bromophenoxide, in boiling dioxane (250 mL). The reaction mixtures were boiled at reflux for 168 h, were concentrated by means of a rotary evaporator, and the concentrate was added to water. The precipitate was collected, washed with ethanol, redissolved in hot tetrahydrofuran [polymer VII (with 56% *p*-bromophenoxy groups) was not soluble in THF and so hot dioxane was used as the solvent for this compound], and precipitated into water. The reprecipitation procedure was carried out once from THF into water, from THF into ethanol, and from THF into pentane. The yields of the polymers represented by VII exceeded 50%. The microanalytical data obtained for these polymers are listed in Table V.

The preparation of polymer III was carried out in the following way. A sodium solution of *p*-bromophenoxide in dioxane, prepared by the method described in the preceding section, was heated to reflux and filtered into a reaction vessel (3-necked, 1-L capacity). A solution of poly(dichlorophosphazene) in dioxane (150 mL) was added dropwise to the sodium *p*-bromophenoxide solution at



reflux. The reaction mixture was heated at reflux for 168 h, concentrated by means of a rotary evaporator, and the concentrate added to water. The precipitate was washed with ethanol and an unsuccessful attempt was made to dissolve the polymer in tetrahydrofuran. The polymer was recovered from tetrahydrofuran by filtration and was then dissolved in boiling dioxane. The hot solution was filtered and reprecipitated into water. The recovered polymer was redissolved in boiling dioxane and was reprecipitated once again into water. The polymer was further reprecipitated from dioxane into ethanol, and then from dioxane into pentane. The yield of III was 30%. The micro-analytical results obtained for III are listed in Table V.

Preparation of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_n$ <sup>18</sup> A solution of phenol (33 g, 0.35 mol) in dioxane (100 mL) was added dropwise to a stirred suspension of sodium hydride (17 g, as a dispersion in oil) and dioxane (100 mL). The reaction mixture was boiled at reflux in order to ensure complete reaction. The hot reaction mixture was filtered and the filtrate was added to a reaction vessel (3-necked, 1-L capacity). To this boiling solution was added dropwise a solution of poly(dichlorophosphazene) (4.0 g, 0.035 mol) in dioxane (200 mL). The reaction mixture was then boiled at reflux for 168 h, concentrated by means of a rotary evaporator, and added to water. The precipitate was collected, washed with ethanol, redissolved in hot THF, and the polymer was reprecipitated into water once again. Reprecipitations were carried out from THF into water for a third time and from THF into ethanol. The yield of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_n$  was over 50%. The microanalytical data obtained for this polymer data are listed in Table V.

General Experimental Procedure for the Preparation of Polymers with p-Lithiophenoxy Substituent Groups. Polymers with p-lithiophenoxy substituent groups were prepared by the reactions of compounds III or VII with n-butyllithium.

Specific experimental details are listed in Table VI. A general procedure for the metal-halogen exchange reactions with VII or III is as follows. The polymers VII (with 5%, 20%, and 35% *p*-bromophenoxy groups) were dissolved in THF (200 mL). Polymers III and VII (56% bromophenoxy groups) required a two-step procedure that involved the use of boiling dioxane, removal of the dioxane by means of a rotary evaporator and then treatment with THF (200 mL). When this procedure was employed, THF-soluble forms of III and VII (56% bromophenoxy) could be obtained. The solutions of III or VII in THF were cooled to  $-40^{\circ}\text{C}$  to  $-60^{\circ}\text{C}$  by means of a Dry Ice-acetone bath (before this step, the reaction vessels were fitted with a rubber septum). *n*-Butyllithium (4 to 6 mL) was then added to the polymer solutions by means of a syringe. This resulted in the immediate formation of yellow or green colored reaction mixtures. Reaction times with *n*-butyllithium were in the range of 0.5 h to 6 h (the most commonly employed reaction time was 0.5 h). The lithiated derivatives were deactivated by treatment with water, diphenylchlorophosphine, carbon dioxide, triphenyltin chloride, or triphenylphosphinegold(I) chloride. These reaction procedures are discussed in the next sections.

General Experimental Procedure for the Deactivation of *p*-Lithiophenoxy Substituent Groups with Water. The lithiated derivatives that were prepared by metal-halogen exchange reactions of *n*-butyllithium with polymers III or VII (5%, 20%, 35%, 56% *p*-bromophenoxy groups) were characterized after their deactivation with water. The specific experimental details involved in these reaction sequences are listed in Table VI. The deactivations of the lithiated intermediates were carried out by the addition of water (by means of a syringe) to the reaction mixtures at  $-40^{\circ}\text{C}$  to  $-60^{\circ}\text{C}$ . After the addition of water, the reaction mixtures were allowed to warm to  $25^{\circ}\text{C}$ . The polymeric products

were isolated by filtration and were purified by reprecipitations from THF into water, from THF into ethanol, and from THF into pentane. The micro-analytical data obtained for the polymers following water deactivation of the lithiated intermediates IV or VIII are listed in Table V.

General Experimental Procedure for the Reactions of p-Lithiophenoxy

Substituent Groups with Diphenylchlorophosphine. Metal-halogen exchange reactions were carried out on polymers III or VII (5%, 20%, 35%, 56% p-bromophenoxy groups) and the resultant lithiated intermediates were deactivated with diphenylchlorophosphine (see Table VII for specific details). Diphenylchlorophosphine was added by means of a syringe to the reaction mixture at  $-40^{\circ}\text{C}$  to  $-60^{\circ}\text{C}$ . The low reaction temperatures were maintained for 3 h, and the reaction mixture was then allowed to warm to  $25^{\circ}\text{C}$ . Ethanol (100 mL) was added to the reaction mixture which was then concentrated by means of a rotary evaporator until the polymer precipitated from solution. The polymeric precipitate was collected by filtration, washed with ethanol, dissolved in THF, and precipitated into ethanol. Reprecipitation was carried out from THF into ethanol for a second time, and then from THF into pentane. Microanalytical data and Tg data obtained for the series of polymers prepared by this general method are listed in Table V. The  $^{31}\text{P}$  NMR spectra of the polymers that possessed pendent phosphine residues consisted of resonances at  $-20$  ppm (phosphazene phosphorus),  $-8$  ppm (phosphine phosphorus) and a small resonance at  $+22$  to  $+24$  ppm (assigned to pendent butylphosphonium bromide residues). These spectra were referenced to an external  $\text{H}_3\text{PO}_4$  sample and employed a  $\text{D}_2\text{O}$  capillary lock. The integrated peak area for these polymers were used as a method for the determination of polymer compositions.

Synthesis of  $[\text{NP}(\text{OC}_6\text{H}_4\text{Sn}(\text{C}_6\text{H}_5)_3\text{-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$ , X. A sample of the compound,  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$ , VII (35% bromophenoxy groups) (0.5 g, 0.0017 mol) was dissolved in THF (350 mL). This solution was then cooled to  $-60^\circ\text{C}$  and *n*-butyllithium (5 mL, 0.008 mol) was added. The reaction mixture was stirred for 1 h at  $-60^\circ\text{C}$ , and then a solution of triphenyltin chloride (6.5 g, 0.017 mol) in THF (75 mL) was added. The reaction mixture was stirred at  $-60^\circ\text{C}$  for 3 h and it was then allowed to warm to  $25^\circ\text{C}$  and was stirred for 14 h. The reaction mixture was concentrated by means of a rotary evaporator and was reprecipitated into a dilute aqueous HCl solution. The polymeric product was collected by filtration and was then reprecipitated from THF into water and from THF into *n*-pentane. The product showed a low solubility in THF. Micro-analytical data were consistent with the formula  $[\text{NP}(\text{OC}_6\text{H}_4\text{Sn}(\text{C}_6\text{H}_5)_3\text{-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$  (see Table V).

Synthesis of  $[\text{NP}(\text{OC}_6\text{H}_4\text{AuP}(\text{C}_6\text{H}_5)_3\text{-P})_{0.52}(\text{OC}_6\text{H}_5)_{1.48}]_n$ , XI. A sample of the compound  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$  (VII) (35% bromophenoxy groups), (0.5 g, 0.0017 mol) was dissolved in THF (200 mL). This solution was then cooled to  $-60^\circ\text{C}$  and *n*-butyllithium (0.6 mL, 0.0010 mol) was added. The reaction mixture was stirred for 1 h at  $-60^\circ\text{C}$ , and a solution of triphenylphosphinegold(I)-chloride (0.50 g, 0.0010 mol) in 25 mL THF was then added. The reaction mixture was stirred for 3 h at  $-60^\circ\text{C}$  and was allowed to warm to  $25^\circ\text{C}$  and was stirred for 1 h. Water (0.5 g, 0.028 mol) was added, the reaction mixture was concentrated by means of a rotary evaporator, and the mixture was filtered. The filtrate was added to *n*-pentane and the precipitate was collected. A  $^{31}\text{P}$  NMR spectrum of the product showed two resonances at +33 ppm and -20 ppm assigned to phosphine and phosphazene phosphorus atoms, respectively. The integrated areas for these resonances were consistent with the formula

$[\text{NP}(\text{OC}_6\text{H}_4\text{AuP}(\text{C}_6\text{H}_5)_3\text{-p})_{0.52}(\text{OC}_6\text{H}_5)_{1.48}]_n$ . Solutions of XI in THF showed signs of decomposition (purple color formation) after only 3 h in air. The compounds  $\text{C}_6\text{H}_5\text{AuP}(\text{C}_6\text{H}_5)_3$  and the cyclic phosphazene,  $[\text{NP}(\text{OC}_6\text{H}_4\text{AuP}(\text{C}_6\text{H}_5)_3\text{-p})_2]_3$ , showed no decomposition in the solid state (these compounds were only slightly soluble in THF. Hence, their solution stability could not be evaluated).

Synthesis of  $[\text{NP}(\text{OC}_6\text{H}_4\text{COOH-p})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$ , XII. A sample of the compound,  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$  (VII) (35% bromophenoxy groups), (0.5 g 0.0017 mol) was dissolved in THF (300 mL). This solution was cooled to  $-60^\circ\text{C}$  and *n*-butyllithium (5 mL, 0.008 mol) was then added. The reaction mixture was stirred for 1 h at  $-60^\circ\text{C}$  and was added to a large excess of Dry Ice in a nitrogen-filled glove bag. The reaction mixture was allowed to warm to  $25^\circ\text{C}$  (in a nitrogen atmosphere), and the white polymeric product was collected by filtration. This was then stirred in an aqueous solution of HCl for 20 min and was once again collected by filtration. This compound proved to be insoluble in THF, boiling dioxane, ethanol, water (pH = 7 or 12), or a THF solution that contained 20% or 50% (by volume) of triethylamine. The infrared spectrum of the polymeric product suggested the presence of carboxylic acid functions (OH stretch at  $3300\text{ cm}^{-1}$  and C=O stretch at  $1750\text{ cm}^{-1}$ ). Microanalysis of this compound was consistent with the formula  $[\text{NP}(\text{OC}_6\text{H}_4\text{COOH-p})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$ , (see Table V).

In a similar manner, the derivative,  $[\text{NP}(\text{OC}_6\text{H}_4\text{COOH-p})_{1.42}(\text{OC}_6\text{H}_4\text{Br-p})_{0.58}]_n$ , (Table V) was prepared by the interaction of  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_2]_n$  (III) (0.5 g, 0.0013 mol) with *n*-butyllithium (4 mL, 0.0064 mol) for 6 h, followed by treatment with carbon dioxide.

Synthesis of  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.050}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9)^\oplus\text{Br}^\ominus\text{-p})_{0.65}(\text{OC}_6\text{H}_5)_{1.30}]_n$ . A sample of  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.050}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2\text{-p})_{0.626}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9)^\oplus\text{Br}^\ominus\text{-p})_{0.024}(\text{OC}_6\text{H}_5)_{1.30}]_n$  (0.5 g, 0.0014 mol), prepared from the reaction of compound VII (35% bromophenoxy) with *n*-butyllithium, followed by treatment with diphenylchlorophosphine (see Table VII), was dissolved in THF (200 mL), and an excess of *n*-butyl bromide (2 g, 0.015 mol) was added. The reaction mixture

was stirred for 1 h at 25°C, and ethanol (300 mL) was added. The reaction mixture was evaporated to dryness by means of a rotary evaporator. The polymeric product was purified by reprecipitations from THF into water and from THF into ethanol. The polymer showed only two  $^{31}\text{P}$  NMR resonances at +24 ppm (phosphonium salt phosphorus) and -20 ppm (phosphazene phosphorus). No resonance at -8 ppm (phosphine phosphorus) could be detected from this sample.  $^{31}\text{P}$  NMR peak area integrations were consistent with the polymer structure,  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.050}(\text{OC}_6\text{H}_4(\text{C}_6\text{H}_5)_2(\text{C}_5\text{H}_9))^{\oplus} \text{Br}^{\ominus} \text{-P})_{0.65}(\text{OC}_6\text{H}_5)_{1.30}]_n$ . The results of this experiment support the contention that pendent phosphonium bromide salt formation is a consequence of the reaction of pendent phosphine residues with butyl bromide (generated in the exchange process) during the purification of the polymer reaction mixtures.

Synthesis of  $[\text{N}_3\text{P}_3(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2\text{-P})_3(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9))^{\oplus} \text{Br}^{\ominus} \text{-P})_3]$ .

A sample of the compound,  $[\text{NP}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2\text{-P})_2]_3$ , (0.50 g,  $2.8 \times 10^{-4}$  mol) was dissolved in dry benzene (100 mL), and butyl bromide (2.0 g,  $1.4 \times 10^{-2}$  mol) was added to this solution. The reaction mixture was boiled at reflux for 12 h. The reaction mixture was then evaporated to dryness in a rotary evaporator. The product was isolated by recrystallization from benzene. A  $^{31}\text{P}$  NMR spectrum of product (in THF) consisted of three resonances at +24.3, +7.9, and -6.5. These were assigned to the phosphonium salt, the phosphazene, and the phosphine phosphorus atoms respectively. The integrated areas of the resonances were in the ratio of 1:1:1.

The Reaction of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$  with Phenyllithium. A sample of the compound,  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$ , (0.5 g,  $7.0 \times 10^{-4}$  mol) was dissolved in THF (100 mL). Phenyllithium (1.2 mL,  $7.2 \times 10^{-4}$  mol) was then added to this solution at 25°C. The reaction mixture was stirred for 3 h and isopropanol (2 mL, 0.023 mol) was

then added to deactivate the organolithium reagent. The reaction solvent was removed by means of a rotary evaporator, and water (25 mL) was added. The water-insoluble material was then extracted with diethyl ether (25 mL) and the diethyl ether was then removed by means of a rotary evaporator. The diethyl ether-soluble product was subjected to mass spectrometric examination. Two parent ions were detected. These were derived from the compounds,  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$  ( $m/e = 693$ ) and  $[\text{NP}(\text{OC}_6\text{H}_5)_5(\text{C}_6\text{H}_5)]_3$  ( $m/e = 677$ ).  $^{31}\text{P}$  NMR spectroscopic data of this reaction mixture indicated only one resonance at +8.2 ppm (this is indicative of the presence of  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$ ). Thus, the compound,  $[\text{N}_3\text{P}_3(\text{C}_6\text{H}_5)(\text{OC}_6\text{H}_5)_5]$  was present in small quantities only in the reaction product mixture (probably less than 5%, because the presence of more than 5% would allow for its detection by  $^{31}\text{P}$  NMR spectroscopic methods). The presence of any amount of the compound,  $[\text{N}_3\text{P}_3(\text{C}_6\text{H}_5)(\text{OC}_6\text{H}_5)_5]$ , implies that the displacement of lithium phenoxide from  $[\text{NP}(\text{OC}_6\text{H}_5)_2]_3$ , by phenyllithium is possible and that this displacement reaction is a prospective mechanism for the crosslinking of polymers that possess p-lithiophenoxy and phenoxy substituent groups.

#### Comparison of the Metal-Halogen Exchange Reaction for Poly(p-bromostyrene).

A sample of poly(p-bromostyrene) (0.5 g, 0.0060 mol) was dissolved in tetrahydrofuran (300 mL). This solution was cooled to  $-50^\circ\text{C}$  and *n*-butyllithium (9 mL, 0.0144 mol) was added rapidly. The reaction mixture was stirred for 6 h at  $-50^\circ\text{C}$ , and water (9 g, 0.5 mol) was then added to deactivate the lithiated species. The reaction mixture was allowed to warm to  $25^\circ\text{C}$ , the reaction mixture was concentrated by means of a rotary evaporator, and the concentrate was added to water. The precipitated polymer was collected by filtration, dried and redissolved in tetrahydrofuran, filtered and reprecipitated into ethanol. A reprecipitation from tetrahydrofuran into pentane was also carried out. Elemental

microanalytical data implied that 15.7% *p*-bromostyrene units remained. Calcd. for  $\{(\text{C}_8\text{H}_8)_x(\text{C}_8\text{H}_7\text{Br})_y\}_{n=100}$  ( $x = 15.7$  and  $y = 84.3$ ): C, 82.50; H, 6.74; Br, 10.76. Anal: C, 81.08; H, 7.22; Br, 10.77. By comparison, the metal-halogen exchange for  $[\text{NP}(\text{OC}_6\text{H}_4\text{Br-}p)_2]_n$  (III) resulted in the replacement of approximately 70% of the *p*-bromophenoxy units after the same reaction time. The greater efficiency of the exchange process for poly(*p*-bromostyrene) probably reflects the fact that only one bromine atom is present per repeat unit for this polymer.

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#### References and Notes:

1. H. R. Allcock, R. L. Kugel, and K. J. Valan, Inorg. Chem., **5**, 1709 (1966).
2. H. R. Allcock, Science, **193**, 1214 (1976); Angew. Chem. (Intern. Ed. English), **16**, 147 (1977); "Phosphorus-Nitrogen Compounds," Academic Press, New York (1972).
3. D. P. Tate, J. Polymer Sci., Polymer Symp., **48**, 33 (1974).
4. R. E. Singler, N. S. Schneider, and G. L. Hagnauer, Polymer Eng. Sci., **15**, 321 (1975).
5. T. L. Evans, T. J. Fuller, and H. R. Allcock, J. Am. Chem. Soc., **101**, 242 (1979).
6. H. R. Allcock and T. J. Fuller, unpublished results.
7. For a review of the applications of the metal-halogen exchange reaction to organic synthesis, see: B. J. Wakefield, "The Chemistry of Organolithium Compounds," Pergamon Press, New York, pp. 51-62 (1974).



8. Incomplete lithiation of a styrene resin (that was partially functionalized with p-bromostyrene residues) has also been reported, M. J. Farrall and J. M. J. Fréchet, J. Org. Chem., 41, 3877 (1976).
9. H. R. Allcock, R. L. Kugel, and E. J. Walsh, J. Chem. Soc., Chem. Commun., 1283 (1970).
10. R. H. Grubbs, L. C. Kroll, and E. M. Sweet, J. Macromol. Sci. - Chem., A7 1047 (1973).
11. C. U. Pittman, Jr., L. R. Smith, and R. M. Haves, J. Am. Chem. Soc., 97, 1742 (1974).
12. N. K. Mathur and R. F. Williams, J. Macromol. Sci., Rev. Macromol. Chem., C15, 117 (1976).
13. Chemical shifts that are downfield from  $H_3PO_4$  are assigned positive signs.
14. F. Camps, J. Castells, J. Font, and F. Vela, Tetrahedron Lett., 1715 (1971).
15. H. Molinari, F. Montavari, S. Quici, and P. Tundo, J. Am. Chem. Soc., 101 3920 (1979).
16. Polymer molecular weight determinations were carried out with the use of tetrahydrofuran as the solvent.
17. H. R. Allcock and C. T.-W. Chu, Macromolecules, 12, 551 (1979).
18. Previously reported syntheses of poly[bis(phenoxy)phosphazene]<sup>1</sup> have employed metallic sodium to generate the phenoxide salt; here we report the use of sodium hydride for this purpose.

Table I

## Properties of Poly(p-bromophenoxyphosphazenes)

Product <sup>a</sup>	GPC MW <sup>b</sup>	Tg <sup>c</sup>
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_2]_n$ , (III)	$5 \times 10^5$	+44
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{1.12}(\text{OC}_6\text{H}_5)_{0.88}]_n$ , (VII) (56%)	$5 \times 10^5$	+21
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$ , (VII) (35%)	$5 \times 10^5$	+ 8
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.4}(\text{OC}_6\text{H}_5)_{1.6}]_n$ , (VII) (20%)	$5 \times 10^5$	---
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.1}(\text{OC}_6\text{H}_5)_{1.9}]_n$ , (VII) (5%)	$5 \times 10^5$	+15
$[\text{NP}(\text{OC}_6\text{H}_5)_2]_n$	$5 \times 10^5$	+14

<sup>a</sup>For the series of polymers represented by VII, the percentage bromophenoxy residues are given in parentheses.

<sup>b</sup>GPC molecular weights are relative to polystyrene standards.

<sup>c</sup>Glass transition measurements were obtained with the use of a torsional braid analyzer.

Table II

## Polymers Derived from Metal-Halogen Exchange

Reactions Followed by Treatment of the Lithiated Derivatives with Water

Composition of Starting Polymers	% Bromophenoxy Substituents after Lithiation and Reaction with Water		
	Lithiation Reaction Times		
	(0.5 h)	(2 h)	(6 h)
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.1}(\text{OC}_6\text{H}_5)_{1.9}]_n$	0%	---	---
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.4}(\text{OC}_6\text{H}_5)_{1.6}]_n$	0%	---	---
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$	2.44%	---	2.14%
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_{1.12}(\text{OC}_6\text{H}_5)_{0.88}]_n$	9.29%	---	3.74%
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-p})_2]_3$	49.1%	45.3%	30.9%

Table III

Polymers Derived from the Metal-Halogen Exchange Reaction and  
Subsequent Treatment with Diphenylchlorophosphine

Starting Polymer	Phosphine Substituted Product		
	% $\text{OC}_6\text{H}_4\text{Br}$	% $\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2$ <sup>a</sup>	% $\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2\text{C}_4\text{H}_9\text{Br}^{\oplus\ominus\text{b}}$
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.1}(\text{OC}_6\text{H}_5)_{1.9}]_n$	0%	5%	----
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.11}(\text{OC}_6\text{H}_5)_{1.6}]_n$	0%	20%	----
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n$	2.5%	31.3%	1.2%
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_{1.12}(\text{OC}_6\text{H}_5)_{0.88}]_n$	9.3%	40.3%	6.2%
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br-P})_2]_n$	49.1%	47.3%	3.7%

<sup>a</sup>The values for pendent phosphine groups are derived from elemental microanalyses for bromine and confirmed by  $^{31}\text{P}$  nmr peak area integrations (error in the  $^{31}\text{P}$  nmr technique is  $\pm 3\%$ ).

<sup>b</sup>The values for butylphosphonium salt formation are derived from elemental microanalyses for bromine and confirmed by  $^{31}\text{P}$  nmr peak area integrations.

Table IV  
Reaction Conditions for the Preparation of Poly(p-bromophenoxyphosphazenes)

Polymer Synthesized	(NPCl <sub>2</sub> ) <sub>n</sub> g (mol)	p-Bromophenol		NaH		Phenol		NaH		% Yield
		g (mol)	g (mol)	g	g (mol)	g (mol)	g (mol)	g	g	
[NP(OC <sub>6</sub> H <sub>4</sub> Br-p) <sub>0.1</sub> (OC <sub>6</sub> H <sub>5</sub> ) <sub>1.9</sub> ] <sub>n</sub>	3.8 (0.033)	0.57 (0.0033)	5	80 (0.85)	31	>50%				
[NP(OC <sub>6</sub> H <sub>4</sub> Br-p) <sub>0.4</sub> (OC <sub>6</sub> H <sub>5</sub> ) <sub>1.6</sub> ] <sub>n</sub>	4.0 (0.035)	2.39 (0.014)	5	27 (0.71)	35	>50%				
[NP(OC <sub>6</sub> H <sub>4</sub> Br-p) <sub>0.7</sub> (OC <sub>6</sub> H <sub>5</sub> ) <sub>1.3</sub> ] <sub>n</sub>	4.3 (0.037)	6.41 (0.037)	10	39 (0.42)	22	>50%				
[NP(OC <sub>6</sub> H <sub>4</sub> Br-p) <sub>1.12</sub> (OC <sub>6</sub> H <sub>5</sub> ) <sub>0.88</sub> ] <sub>n</sub>	4.0 (0.035)	8.95 (0.053)	10	40 (0.43)	21	>50%				
[NP(OC <sub>6</sub> H <sub>4</sub> Br-p) <sub>2</sub> ] <sub>n</sub>	3.0 (0.026)	9.0 (0.52)	25	---	--	>50%				

Table V

## Table of Analytical Results

Polymer Structure	% C	% H	% Br
$[\text{NP}(\text{OC}_6\text{H}_5)_2]_n^a$	Calcd. 62.34	4.19	---
	Anal. 60.15	4.33	---
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.1}(\text{OC}_6\text{H}_5)_{1.9}]_n^{a,b}$ VII (5%)	Calcd. 60.47	4.16	3.04
	Anal. 60.38	4.35	3.04
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.4}(\text{OC}_6\text{H}_5)_{1.6}]_n$ VII (20%)	Calcd. 54.84	3.66	12.18
	Anal. 54.16	4.01	12.16
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n^a$ VII (35%)	Calcd. 50.33	3.25	19.50
	Anal. 51.03	3.71	19.45
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{1.12}(\text{OC}_6\text{H}_5)_{0.88}]_n^a$ VII (56%)	Calcd. 45.09	2.78	28.02
	Anal. 42.40	2.78	28.01
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_2]_n^{a,c}$ III	Calcd. 37.04	2.06	41.10
	Anal. 37.13	2.08	40.96
$[\text{NP}(\text{OC}_6\text{H}_5)_2]_n^d$	Calcd. 62.25	4.32	0.00
	Anal. 56.70	3.99	<0.15

Table V (continued)

Polymer Structure		% C	% H	% Br
$[\text{NP}(\text{OC}_6\text{H}_5)_2]_n^e$	Calcd.	62.25	4.32	0.00
	Anal.	61.38	4.35	<0.15
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.05}(\text{OC}_6\text{H}_5)_{1.95}]_n^f$	Calcd.	61.32	4.24	1.66
	Anal.	61.62	4.24	1.66
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.19}(\text{OC}_6\text{H}_5)_{1.81}]_n^g$	Calcd.	57.87	3.97	6.04
	Anal.	58.62	3.99	6.04
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.98}(\text{OC}_6\text{H}_5)_{1.02}]_n^h$	Calcd.	46.69	2.92	26.42
	Anal.	44.89	2.73	25.33
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.9}(\text{OC}_6\text{H}_5)_{1.1}]_n^i$	Calcd.	47.61	3.01	23.92
	Anal.	25.91	3.21	23.91
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.04}(\text{OC}_6\text{H}_5)_{1.96}]_n^j$	Calcd.	60.73	4.39	1.46
	Anal.	61.44	4.25	1.46
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.07}(\text{OC}_6\text{H}_5)_{1.93}]_n^k$	Calcd.	60.79	4.19	2.52
	Anal.	59.20	3.89	2.52

Table V (continued)

Polymer Structure	% C	% H	% Br
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.62}(\text{OC}_6\text{H}_5)_{1.38}]_n^{\text{l}}$	Calcd. 51.47	3.35	17.65
	Anal. 45.56	2.84	17.61
$[\text{NP}(\text{OC}_6\text{H}_4\text{P}(\text{C}_4\text{H}_5)_2)_{0.1}(\text{OC}_6\text{H}_5)_{1.9}]_n^{\text{m}}$	Calcd. 63.52	4.37	0.00
	Anal. 63.60	4.28	0.17
$[\text{NP}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2)_{0.4}(\text{OC}_6\text{H}_5)_{1.6}]_n^{\text{n}}$	Calcd. 66.08	4.47	0.19
	Anal. 64.54	4.71	0.19
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.05}(\text{OC}_6\text{H}_5)_{1.3}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_{0.63}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9)^+\text{Br}^-)_{0.024}]_n^{\text{o}}$	Calcd. 66.73	4.06	1.62
	Anal. 57.96	4.48	1.61
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.19}(\text{OC}_6\text{H}_5)_{0.58}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2)_{0.81}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9)^+\text{Br}^-)_{0.12}]_n^{\text{p}}$	Calcd. 65.50	4.45	5.66
	Anal. 62.43	4.15	5.67
$[\text{NP}(\text{OC}_6\text{H}_4\text{Br})_{0.98}(\text{OC}_6\text{H}_9\text{P}(\text{C}_6\text{H}_5)_2)_{0.95}(\text{OC}_6\text{H}_4\text{P}(\text{C}_6\text{H}_5)_2(\text{C}_4\text{H}_9)^+\text{Br}^-)_{0.07}]_n^{\text{q}}$	Calcd. 58.16	3.73	16.64
	Anal. 56.61	3.68	16.65
$[\text{NP}(\text{OC}_6\text{H}_4\text{COOH})_{0.7}(\text{OC}_6\text{H}_5)_{1.3}]_n^{\text{r}}$	Calcd. 58.17	3.83	---
	Anal. 57.86	3.77	0.2
$[\text{NP}(\text{OC}_6\text{H}_4\text{COOH})_{1.42}(\text{OC}_6\text{H}_4\text{Br})_{0.58}]_n^{\text{s}}$	Calcd. ---	---	13.68
	Anal. ---	---	13.86



Table V (continued)

Polymer Structure	% C	% H	% Br
[NP(OC <sub>6</sub> H <sub>4</sub> Sn(C <sub>6</sub> H <sub>5</sub> ) <sub>3</sub> ) <sub>0.7</sub> (OC <sub>6</sub> H <sub>5</sub> ) <sub>1.3</sub> ] <sub>n</sub> <sup>t</sup>	Calcd.	62.14	4.17
	Anal.	60.12	3.80
			0.3

<sup>a</sup>Percentage of chlorine in samples: [NP(OC<sub>6</sub>H<sub>5</sub>)<sub>2</sub>]<sub>n</sub>, 0.30%; VII (5%), 0.19%; VII (20%), 0.20%; VII (35%), 0.28; VII (56%), 0.27%; III, 0.046%.

<sup>b</sup>Calcd.: P, 12.73; N, 5.70. Anal.: P, 13.70; N, 5.88.

<sup>c</sup>Calcd.: P, 7.97; N, 3.60. Anal.: P, 7.87; N, 3.46.

<sup>d</sup>From lithiation of VII (5%) (0.5 h) and water quench.

<sup>e</sup>From lithiation of VII (20%) (0.5 h) and water quench.

<sup>f</sup>From lithiation of VII (35%) (0.5 h) and water quench.

<sup>g</sup>From lithiation of VII (56%) (0.5 h) and water quench.

<sup>h</sup>From lithiation of III (0.5 h) and water quench.

<sup>i</sup>From lithiation of III (2 h) and water quench.

<sup>j</sup>From lithiation of VII (35%) (6 h) and water quench.

<sup>k</sup>From lithiation of VII (56%) (6 h) and water quench.

<sup>l</sup>From lithiation of III (6 h) and water quench.

Table V (continued)

- <sup>m</sup>From lithiation of VII (5%) (0.5 h) and diphenylchlorophosphine quench.
- <sup>n</sup>From lithiation of VII (20%) (0.5 h) and diphenylchlorophosphine quench.
- <sup>o</sup>From lithiation of VII (35%) (0.5 h) and diphenylchlorophosphine quench.
- <sup>p</sup>From lithiation of VII (56%) (0.5 h) and diphenylchlorophosphine quench.
- <sup>q</sup>From lithiation of III (0.5 h) and diphenylchlorophosphine quench.
- <sup>r</sup>From lithiation of VII (35%) (0.5 h) and CO<sub>2</sub> quench and then acid treatment.  
Oxygen analysis: calculated, 20.78%; analyzed, 22.34%.
- <sup>s</sup>From lithiation of III (6 h) and CO<sub>2</sub> quench and then acid treatment. Oxygen  
analysis: calculated, 22.82%; analyzed 22.76%.
- <sup>t</sup>From lithiation of VII (35%) (0.5 h) and triphenyltinchloride.

Table VI  
Reaction Conditions for Water Deactivation  
of Poly(p-lithiophenoxyphosphazenes)

Starting Polymer <sup>a</sup> g (mol)	<u>n</u> -BuLi mL (mol)	Reaction Time (h)	Water mL (mol)
VII (5%) 0.50 (0.0021)	4 (0.0064)	0.5	4 (0.22)
VII (20%) 0.50 (0.0019)	4 (0.0064)	0.5	4 (0.22)
VII (35%) 0.58 (0.0020)	4 (0.0064)	0.5	4 (0.22)
VII (35%) 0.58 (0.0020)	4 (0.0064)	6	4 (0.22)
VII (56%) 0.58 (0.0018)	5 (0.008)	0.5	5 (0.28)
VII (56%) 0.54 (0.0017)	5 (0.008)	6	5 (0.28)
III 0.55 (0.0014)	6 (0.0096)	0.5	6 (0.33)
III 0.62 (0.0016)	6 (0.0096)	2	6 (0.32)
III 0.54 (0.0014)	6 (0.0096)	6	6 (0.33)

<sup>a</sup>% Bromophenoxy substituent groups in parentheses.

Table VII

Conditions for the Reaction of  
Poly(p-lithiophenoxyphosphazenes) with Diphenylchlorophosphine

Starting Polymer <sup>a</sup> mL (mol)	n-C <sub>4</sub> H <sub>9</sub> Li mL (mol)	Reaction Time (h)	ClP(C <sub>6</sub> H <sub>5</sub> ) <sub>2</sub> mL (mol)	Product <sup>b</sup>
VII (5%) 0.50 (0.0021)	4 (0.0064)	0.5	4 (0.018)	c
VII (20%) 0.50 (0.0019)	4 (0.0064)	0.5	6 (0.027)	d
VII (35%) 0.57 (0.0020)	4 (0.0064)	0.5	8 (0.036)	e
VII (56%) 0.61 (0.001)	5 (0.0080)	0.5	10 (0.045)	f
III 0.62 (0.0016)	6 (0.0096)	0.5	12 (0.054)	g

<sup>a</sup>% Bromophenoxy substituent groups in parentheses.

<sup>b</sup>Product compositions were determined from elemental microanalytical data obtained for bromine. These calculations are based on the assumption that the metal-halogen exchange reaction proceeds to the same point as determined for the proton quenching experiments (see Table VII-6). <sup>31</sup>P NMR data for this phosphine series implied that this assumption is accurate.

<sup>c</sup>[NP(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>-P)<sub>0.1</sub>(OC<sub>6</sub>H<sub>5</sub>)<sub>1.9</sub>]<sub>n</sub>: T<sub>g</sub> = +8°C.

<sup>d</sup>[NP(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>-P)<sub>0.4</sub>(OC<sub>6</sub>H<sub>5</sub>)<sub>1.6</sub>]<sub>n</sub>.

<sup>e</sup>[NP(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>-P)<sub>0.63</sub>(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>(C<sub>4</sub>H<sub>9</sub>)<sup>⊕</sup>Br<sup>⊖</sup>-P)<sub>0.02</sub>(OC<sub>6</sub>H<sub>4</sub>Br-P)<sub>0.05</sub>-  
(OC<sub>6</sub>H<sub>5</sub>)<sub>1.3</sub>]<sub>n</sub>: T<sub>g</sub> = +17°C.

<sup>f</sup>[NP(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>-P)<sub>0.81</sub>(OC<sub>6</sub>H<sub>4</sub>P(C<sub>4</sub>H<sub>9</sub>)<sup>⊕</sup>Br<sup>⊖</sup>-P)<sub>0.12</sub>(OC<sub>6</sub>H<sub>4</sub>Br-P)<sub>0.19</sub>-  
(OC<sub>6</sub>H<sub>5</sub>)<sub>0.88</sub>]<sub>n</sub>.

<sup>g</sup>[NP(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>-P)<sub>0.95</sub>(OC<sub>6</sub>H<sub>4</sub>P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub><sup>⊕</sup>Br<sup>⊖</sup>-P)<sub>0.07</sub>(OC<sub>6</sub>H<sub>4</sub>Br-P)<sub>0.98</sub>]<sub>n</sub>:  
T<sub>g</sub> = +35°C.

Legend to Figure 1

A plot of the percent remaining *p*-bromophenoxy pendent groups versus reaction time following the reaction with excess *n*-butyllithium for polymers that possessed initially 100% *p*-bromophenoxy substituent groups (polymer III, curve A) and 56% *p*-bromophenoxy, 44% phenoxy substituent groups (polymer VII, curve B). The percentages of residual *p*-bromophenoxy groups were obtained by elemental microanalytical determinations for bromine after deactivation of the lithiated derivatives with water and conversion of lithioaryl groups to aryl. The curves illustrate the dramatic decrease in lithiation rate after approximately 50% of the pendent groups have reacted.

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